Steps to Drawing a Bohr Diagram

1) Find the element you're drawing a Bohr Diagram for on the Periodic Table
2) Determine the element's atomic number (from the element's box on the periodic table).
3) The atomic number is the same as the number of electrons that element has. That's how many electrons you need to draw.
4) Determine how many shells (or energy levels) the element's atom will have: Elements in the first period will have one shell, in the second period, two shells, and so forth.
5) Draw a nucleus with your element symbol in the middle.
6) Draw the appropriate number of shells around the nucleus, as circles
7) Place the electrons on the shells, noting how many electrons each shell can hold. See table below. The formula is 2n^2 (2 n squared, where n is the shell number)

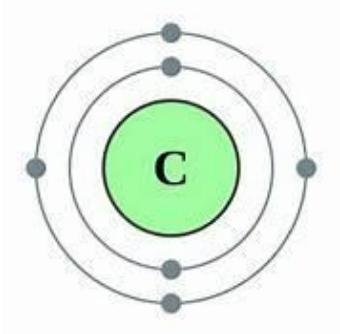
Shell Number	Maximum Number of Electrons
1	2
2	8
3	18
4	32

Example 1: Carbon

Atomic number 6 (so carbon has 6 electrons)

Carbon has 2 shells because it is in period 2

- First shell will have 2 electrons
- Second shell will have the remaining (Carbon has 4 valence electrons)



Example 2: Aluminum

Atomic number 13 (so Aluminum has 13 electrons)

Aluminum has 3 shells because it's in period 3

- First shells will have 2 electrons
- Second shell will have 8 electrons
- Third shell will have the remaining 3 electrons (The 3 electrons in the outer shell are the valence)